Unit 2: Lesson 1 Vocabulary (Text 2.1-2.2)

Name _____ period _____ date _____

Atoms	The basic unit of matter; smallest part of an element
Atomic nucleus	The center of the atom; contains all of the atom's mass
Proton	Positively charged subatomic particle found in the nucleus
Neutron	Found in nucleus; has no charge
Electron	Negatively charged subatomic particle found in energy level surrounding the nucleus
Atomic Number	The number of protons (and electrons) in an atom. The identification of the atom
Atomic Mass	The amount of protons plus the amount of neutrons
Element	A pure substance made up of only one type of atom.
Isotope	Atoms of the same element that have different atomic masses (different amount of neutrons)
Radioactive isotope	Have an unstable nucleus that will release particles and energy
Chemical compound	The chemical combination of two or more elements in definite proportions
Molecule	The smallest part of a compound
Chemical bond	Forces of attraction that hold atoms together in compounds. Form when electrons from atoms interact
Covalent bond	Form when atoms share electrons. Strong, high-energy bonds
Ionic bond	Form when electrons are transferred from one atom to another. A bond between ions.
Ion	Form when atoms gain (- charge) or lose (+ charge) electrons
Polar molecule	A molecule that has a + end and a – end. Most ionic compounds and some covalent compounds
Hydrogen bond	Forms between hydrogen and a – charged atom. Weak bonds that give molecules their shape
Adhesion	Attraction to different substances
Capillarity	Spreading through fine pores
Cohesion	Attraction to like substances
Specific heat	The amount of energy needed to change the temperature of a substance. Heat of vaporization and heat of fusion.
Density	Mass divided by volume. The amount of space an object takes up.